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e. electrolysis of molten potassium chloride decomposition;  $2\text{KCl(l)} \rightarrow 2\text{K(s)} + \text{Cl}_2\text{(g)}$  4.) (l 2 4((( ())) (((

### 8 Chemical Equations and Reactions

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### Chapter 8 Mastery Worksheet Section 6 Supply And Demand ...

8 (propane gas) Each carbon is  $\frac{8}{3}$  and each hydrogen is 1. 8. Assign oxidation numbers to each element in the compounds found in the following situations: a. Rust,  $\text{Fe}_2\text{O}_3$ , forms on an old nail. Each iron is 3 and each oxygen is 2. b. Nitrogen dioxide,  $\text{NO}_2$ , pollutes the air as a component of smog. Nitrogen is 4 and each oxygen is 2.

### 7 Chemical Formulas and Chemical Compounds

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## mc06se cFMsr i-vi

Prentice Hall Chemistry Worksheets. Chapter 22 Hydrocarbon Compounds 237 SECTION 22.1 HYDROCARBONS (pages 693-701) This section describes the bonding in hydrocarbons and distinguishes straight-chain ...

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Example:  $(8.97 \times 10^4) - (2.62 \times 10^3) = (8.97 \times 10^4) - (0.262 \times 10^4) = 8.71 \times 10^4$ ; For multiplication problems you can multiply the coefficients and then add the exponents on your powers of ten. The final answer must be adjusted so that the coefficient is equal to or greater than 1 and less than 10. For example,

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